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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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Use the chemical formula to find the number of atoms in one molecule and multiply this number by Avogadro's number, the number of particles in one mole. atom  $6.02 \cdot 10^{23}$  O  $2 \cdot 6.02 \cdot 10^{23}$  ion Na<sup>+</sup>  $6.02 \cdot 10^{23}$  formula unit NaCl  $6.02 \cdot 10^{23}$   $6.02 \cdot 10^{23}$  representative particles of a substance molecule formula unit atom

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